

Sustained Release of Naltrexone from Poly(*N*-Isopropylacrylamide) Microgels

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Received 12 September 2013; revised 17 October 2013; accepted 18 October 2013

Published online 11 November 2013 in Wiley Online Library (wileyonlinelibrary.com). DOI 10.1002/jps.23780

ABSTRACT: The release of the opioid antagonist naltrexone from neutral poly(*N*-isopropylacrylamide) (PNIPAAM) microgels and negatively charged PNIPAAM microgels containing acrylic acid groups (PNIPAAM-*co*-PAA) has been studied at various microgel and drug concentrations. The release curves were found to be well represented by the Weibull equation. The release rates were observed to be dependent on the microgel concentration. At most conditions, the release from the charged microgels was slower than for the neutral microgels. In addition, the charged microgels exhibited a release lag time, which was dependent on the microgel concentration. No significant lag time could be observed for the neutral microgels. Increasing the naltrexone concentration did not significantly affect the release rates from the neutral microgels, but the release from the charged microgels became faster. The microgels did not exhibit any significant cytotoxic effect on HeLa cells at the tested concentrations. © 2013 Wiley Periodicals, Inc. and the American Pharmacists Association *J Pharm Sci* 103:227–234, 2014

Keywords: controlled release; drug delivery systems; macromolecular drug delivery; microencapsulation; microparticles; polymeric drug delivery systems

INTRODUCTION

Naltrexone (Fig. 1) is an opioid antagonist, and has been used to treat heroin and alcohol addiction. However, oral naltrexone formulations have drawbacks such as patient noncompliance,¹ fluctuating plasma levels,² and the extensive first-pass metabolism of naltrexone.³ Injectable depot formulations that can give a sustained release of naltrexone over an extended period of time could prevent the problems encountered with oral dosage forms. Several systems such as poly(L-lactide) microspheres,^{4,5} poly lactide-*co*-glycolide (PLGA) microspheres,⁶ poly(D,L-lactide-*co*-glycolide) microspheres,⁷ polyethylene glycol-*graft*-methyl methacrylate crosslinked nanoparticles,⁸ and nanoparticles produced from a blend of poly(*N*-isopropylacrylamide-acrylamide-vinylpyrrolidone) and PLGA⁹ have been suggested for sustained release dosage forms for naltrexone. There is however still need for improving the release profile and drug loading capacity of the systems. Utilizing new kinds of thermosensitive *in situ* gelling microgels might therefore be of interest.

The release of a drug that is encapsulated into micro- or nanoparticles is dependent on a number of factors such as particle size,^{10–13} drug size,^{14,15} the degree of swelling of the particles,^{16–18} erosion of the particles,¹⁹ and associative interactions between the drug and the particles, for example, hy-

drophobic interactions,¹⁵ hydrogen bonds,^{20,21} or electrostatic interactions.¹⁴

Nano- and microparticles that are formed using thermosensitive polymers with a lower critical solution temperature (LCST) have a swollen structure at temperatures below the transition temperature and a more compact (collapsed) structure at higher temperatures. This effect can be utilized both to increase the loading capacity and modulate the release of a drug from the particles. Increased loading rate and capacity may be expected for drug loading in the swollen state, whereas the drug release is often found to be slower when the systems are heated to temperatures above the LCST of the particles due to reduced porosity. This phenomenon has been observed for several different systems such as: core-shell nanoparticles containing a poly(L-lactic acid) core and a poly(*N*-isopropylacrylamide) (PNIPAAM) thermosensitive shell,²² alginate-hydroxypropylcellulose microbeads,²³ poly(*N*-vinyl caprolactam) nanoparticles,²⁴ poly(*N*-isopropylacrylamide-*co*-caprolactam) microspheres,²⁵ and PNIPAAM microspheres grafted with poly(itaconic acid) groups.¹⁴ Interestingly, microspheres of elastin-like polypeptides (ELP) exhibited faster release rates above the LCST due to the opening of micropores in the microspheres by the contraction of the thermosensitive ELP molecules.²⁶ Faster release rates above the LCST was also observed for chitosan-*g*-poly(*N*-vinylcaprolactam) and chitosan-*g*-PNIPAAM nanoparticles.^{20,21} This was explained by the ability of the drug to form hydrogen bonds with the nanoparticles at temperatures below the LCST. This capacity was lost above the LCST, and hence faster release rates are promoted.

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Journal of Pharmaceutical Sciences, Vol. 103, 227–234 (2014)

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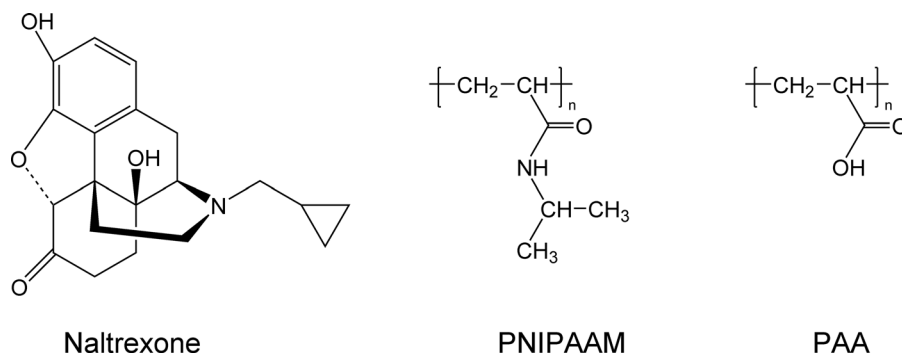


Figure 1. Chemical structure of naltrexone, PNIPAAm, and PAA.

For poly(*N*-isopropylacrylamide-*co*-acrylamide) and poly(*N*-isopropylacrylamide-*co*-*N*-hydroxymethyl acrylamide) microspheres, it was found that even though the drug release is slower above the LCST, there is an enhanced release of drug for a short period when the temperature is increased from below to above the LCST.^{15,27} This effect was thought to be due to drug molecules that are dissolved in the solvent inside the swollen microspheres. When the samples are heated up, these drug molecules are squeezed out together with the solvent during the collapse of the network. Similar effects have also been observed for other thermoresponsive systems.²⁸

In this study, we have investigated PNIPAAm microgels as a sustained drug delivery vehicle for naltrexone. The microgel concentrations utilized in this study are quite high, and it is not possible to measure their sizes using light scattering methods due to multiple scattering. However, previous studies of the same microgels at lower concentrations show that the microgels are swollen and have an open structure at low temperatures.²⁹ As the temperature is increased the microgels shrink, and at temperatures above the LCST of PNIPAAm (ca. 32°C) large aggregates with a compact structure are formed.

The drug is loaded into the microgels at a low temperature. At these conditions, the microgels are swollen in the solvent, and this open structure should make it easier for the drug to diffuse into the microgels. The microgel suspensions have a low viscosity and can be easily administered to the patient through injection. When the samples are heated up to body temperature, the microgels collapse and they also aggregate into larger structures.²⁹ A schematic illustration of a possible

mechanism for loading of naltrexone into the PNIPAAm microgels is depicted in Figure 2. The conjecture is that both the more compact structure and the formation of larger aggregates at 37°C should slow down the diffusion of naltrexone out of the microgels. As naltrexone contains hydrophobic domains, associative interactions between the drug and the PNIPAAm microgels could also result in slower drug release rates.

Naltrexone is positively charged at physiological pH.³⁰ We have therefore also studied PNIPAAm microgels containing 1 mol % acrylic acid groups to see if the electrostatic interactions between the drug and the oppositely charged microgels can slow down the release rates. At low microgel concentrations, this low charge density was observed to have little effect on the contraction and aggregation behavior of the microgels in water.²⁹ Any changes in the observed release profiles are therefore mainly expected to be due to electrostatic interactions between the drug and the microgels.

As for any proposed drug delivery system, the toxicity of the formulation is of vital importance. We have therefore also conducted *in vitro* cytotoxicity studies on the microgels at the considered concentrations.

MATERIALS AND METHODS

The naltrexone (USPH1453504) was purchased from VWR International (West Chester, Pennsylvania, USA), and the chemicals used for making the buffer solution were purchased from Sigma–Aldrich (St. Louis, Missouri, USA).

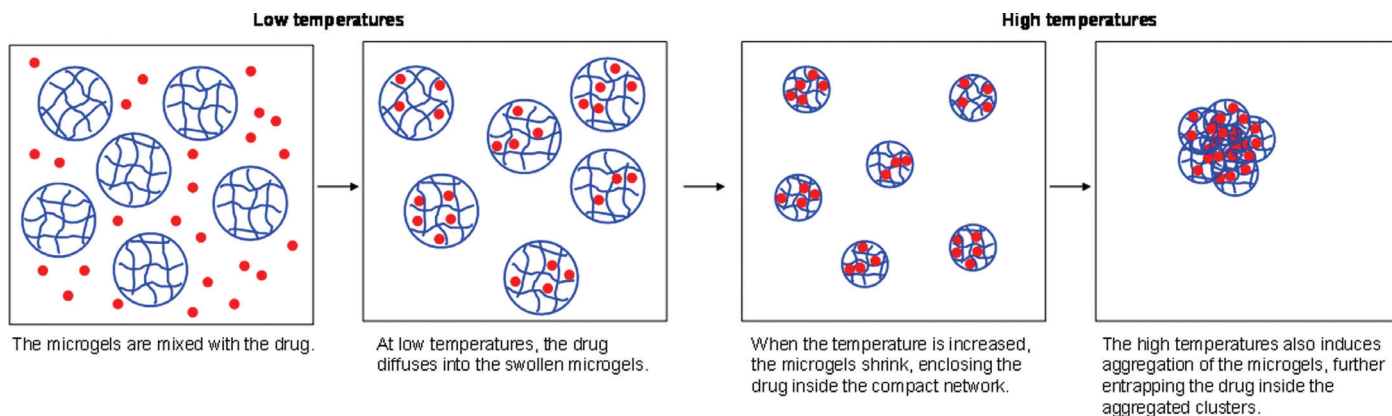


Figure 2. Schematic illustration of the proposed loading of naltrexone into the microgels.

Preparation of PNIPAAm Microgels

Negatively charged PNIPAAm microgels (MG-04) containing 1 mol % acrylic acid groups (PNIPAAm-co-PAA) and neutral PNIPAAm microgels (without acrylic acid groups; MG-06) were synthesized as described previously.²⁹ The structures of PNIPAAm and PAA are shown in Figure 1.

Sample Preparation

A PBS was used for the preparation of the samples for the dissolution tests. The buffer contained 16.8 mM Na₂HPO₄, 1.4 mM KH₂PO₄, 137 mM NaCl, and 3.1 mM sodium azide. The pH of the buffer was adjusted with acetic acid to 7.4.

For the release experiments, the microgels were dispersed in PBS at the desired concentrations (0–4 wt %), and stirred for at least 24 h. Afterward, the samples were placed in the fridge (ca. 4°C) for at least another 24 h to ensure homogeneous suspensions. Naltrexone (0.02–0.08 wt %) was loaded into the microgels by dissolving it directly in the microgel suspensions. The samples were then stirred for at least 24 h, and then placed in the fridge for at least 48 h before the release experiments were commenced. The concentrations given for naltrexone and the microgel suspensions are of the prepared suspensions before the samples were diluted down by the initiation of the release experiments.

In Vitro Cytotoxicity Assay of the Microgels

HeLa cells (human epithelial cervical carcinoma cell line) were maintained in culture at 37°C, under 5% CO₂, in Dulbecco's modified Eagle's medium–high glucose (DMEM–HG; Sigma), supplemented with 10% (v/v) heat inactivated fetal bovine serum (Sigma), penicillin (100 U/mL), and streptomycin (100 µg/mL). HeLa cells grown in monolayer were detached by treatment with 0.25% trypsin solution (Sigma).

The cell viability in the presence of the microgels was investigated by the modified Alamar Blue assay. This method, which has been extensively used to assess the viability and/or proliferation of mammalian cells, employs a redox indicator whose color (and fluorescence) changes in response to cell metabolic activity, the blue nonfluorescent oxidized form (resazurin) originating a pink and fluorescent molecule (resorufin) upon reduction.³¹ Briefly, HeLa cells were seeded at a cell density of 50,000 cells/well in 48-well plates and incubated at 37°C for 24 h. The microgels were suspended in Opti-MEM (reduced serum) medium at different concentrations and were added to the cells. The plates were left to incubate at 37°C for 48 h. The medium was replaced with 300 µL of 10% (v/v) Alamar Blue (Sigma) dye in DMEM cell culture medium and the plates were further incubated for 45 min at 37°C. An aliquot of the supernatant (150 µL) was collected from each well and transferred to 96-well plates. The absorbance at 570 nm (*A*₅₇₀) and 600 nm (*A*₆₀₀) (information provided by the supplier) was measured in a SPECTRAMax PLUS 384 spectrophotometer. Cell viability was calculated according to the equation:

$$\text{Cell viability (\%)} = \frac{(A_{570} - A_{600}) \text{ of treated cells}}{(A_{570} - A_{600}) \text{ of control cells}} \times 100 \quad (1)$$

Release Experiments

The release experiments were conducted utilizing a SOTAX USP Apparatus four flow through cell dissolution apparatus at

37°C. 2.5 mL of the samples was placed in the flow-through cells (which were filled with 1 mm glass beads to avoid turbulence and ensure a laminar flow of the buffer through the cells³² and left to equilibrate at 37°C for 10 min before the release experiments were initiated. During the release studies, 22.5 mL of PBS (giving a total volume of 25 mL) was continuously pumped through the flow through cells at a flow rate of 4 mL/min. The microgels were prevented from escaping the cells by a combination of two filters: a membrane of regenerated cellulose with a 1 kDa cutoff (Z-MEM-AQU-425N from Postnova) followed by a dialysis membrane (Spectra/Por 6 with a cutoff of 8000 Da). The buffer was circulating in a closed system, and the circuit was connected to a spectrophotometer (Pharmacia LKB – Ultrospec III), which detected the released naltrexone by measuring the absorbance at a wavelength of 282 nm every 2nd min for a total release time of 5 h.

The threshold concentration for departure from sink conditions for naltrexone is approximately 0.5 wt %, ³³ and accordingly all release experiments have been conducted well within sink conditions.

Evaluation of the Release Rates

The release data have been fitted to the Weibull equation^{34,35}:

$$\frac{M_t}{M_\infty} = 1 - \exp \left[- \left(\frac{t - t_0}{\tau_d} \right)^\beta \right] \quad (2)$$

where *M_t* is the amount of drug released at the time *t*, *M_∞* is the maximum amount of drug released at infinite time, *t₀* is the dissolution lag time, the rate constant *τ_d* is the time after *t₀* when 63.2% of the drug has been released, and *β* is the shape parameter of the release curve.

For the neutral microgels, *t₀* was always found to be zero within the experimental errors. The release curves for these microgels were therefore fitted using *t₀* ≡ 0 to reduce the number of fitting parameters.

The release curves for the samples where no lag time was observed were also fitted to an equation where an extra term was added to Eq. (2) to accommodate for a possible burst release effect. However, this resulted in very poor fits, and the burst release doses were always found to be zero within the experimental errors. This indicates that the systems do not exhibit any burst release, and the extra term were therefore not used in the analyses of the data.

The Weibull equation is often found to give a good fit to experimental drug release data over the whole experimental time scale.^{36–40} In spite of its statistical origin, the Weibull model can be used to gain information about the physicochemical aspects behind the release profile. Both experimental data^{10–13} and Monte Carlo simulations^{41,42} show that the drug is released slower from larger particles, that is, *τ_d* becomes higher as the particle size is increased. The reduction of the specific leak surface (surface area/volume) of the particles with increasing size plays an important role in this phenomenon.^{41,42}

The shape parameter, *β*, may be related to the release mechanism.^{36,43} Monte Carlo simulations also suggests that the value of *β* depends both on the size of the particles (*β* decreases when the particles becomes larger), on the shape of the particles, and on the ability of the drug molecules to move inside the particle matrix.⁴² For Fickian diffusion, an increase of *β* should indicate less disorder inside the particles.³⁶

RESULTS AND DISCUSSION

In Vitro Cytotoxicity

Because the microgels are intended for drug delivery purposes, it is important that they do not exhibit any significant cytotoxic effect. It has previously been found that PNIPAAm nanoparticles did not exhibit any cytotoxicity at concentrations up to about 1 wt %, whereas NIPAAm monomers were found to be toxic even at low concentrations.⁴⁴ We have tested the two microgels used in this study on HeLa cells using the same microgel concentrations as for the release experiments. The cell viability for different concentrations of the two microgels is displayed in Figure 3. Comparing the viability of cells incubated with the microgels to that of control cells, taken as 100%, it is noticeable that microgels did not exhibit any significant cytotoxicity even at relatively high concentrations. At the lower concentrations assayed, the neutral PNIPAAm microgels apparently induced an increase of cell viability relative to control cells. To correctly interpret the meaning of this effect, it is necessary to take into account the principle that underlies the technique used to assess cell viability. The Alamar Blue assay uses a dye as an oxidation–reduction indicator, whose reduction occurs inside the cells³¹ and involves reductases and the mitochondrial electron chain.⁴⁵ Therefore, compounds that promote an increase of nonviable cells, which have no metabolic capacity, will induce lower reduction of the dye. On the contrary, agents or conditions that enhance cell metabolic activity and/or promote higher rates of cell growth will induce higher levels of dye reduction. Both these phenomena (improvement of cell metabolic activity and increase of cell proliferation) can be on the basis of the high values of cell viability (>100%) found at low concentrations of neutral PNIPAAm microgels. This effect has been observed before for other systems, and was attributed to the surface properties of the systems.⁴⁶ It is also possible that the microgels act as a scaffold, thereby increasing the proliferation rates of the cells.⁴⁷

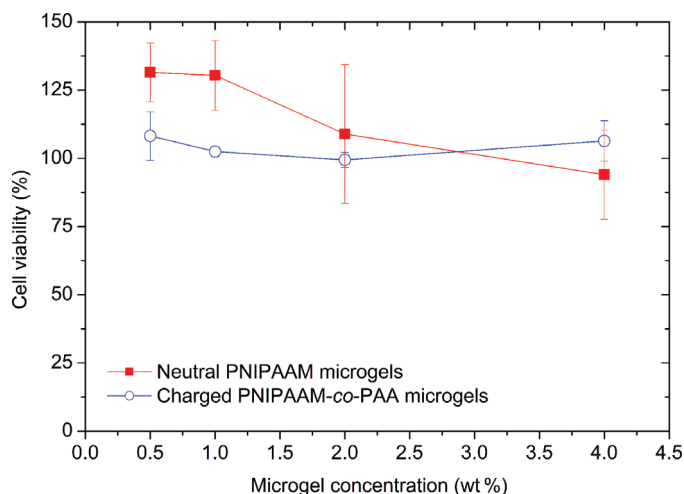


Figure 3. Effects of microgels on the viability of HeLa cells with increasing concentrations at 37°C. Values of cell viability, measured by the Alamar Blue assay as described in *Materials and Methods*, are expressed as a percentage of the untreated cells (control) and are means \pm standard deviation obtained from triplicates in at least three independent experiments.

In Vitro Release

The release profiles of 0.02 wt % naltrexone from suspensions of the two microgels in phosphate buffer (pH 7.4) at 37°C are depicted in Figure 4. When a buffer solution containing naltrexone in the absence of any microgel suspension (0 wt % microgel) is used, the time it takes for the system to detect 100% of the naltrexone is very short (ca. 5 min). This indicates that the detection delay time of the experimental setup is fast. For both microgel systems, the release rate is observed to decrease as the microgel concentration is raised. Interestingly, a release lag time is evident for the charged microgels but not for the neutral microgels. This suggests that the charges modulate the way in which the naltrexone is released from the microgels. To quantify the variations in the release profiles, the curves have been fitted with the Weibull equation (Eq. 2). The solid lines in Figure 4 represent the fitted curves, and this illustrates that the data is well represented by the Weibull equation. The good fitting of the experimental data with the Weibull equation suggests that practically all of the naltrexone is encapsulated within the particles leaving an insignificant amount of unencapsulated naltrexone in the suspension. Accordingly, naltrexone is probably attached to the microgels through associative interactions, as drug molecules that are just dissolved in the solvent inside the swollen microgels are expected to be pushed out of the microgels when they are heated up above the LCST.^{15,27}

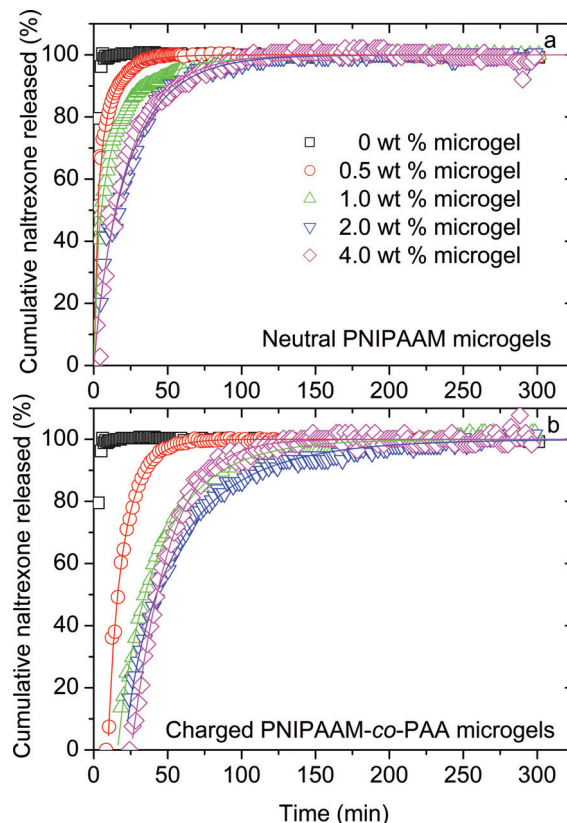


Figure 4. *In vitro* release profiles of naltrexone (0.02 wt %) from the two microgels in buffered suspensions (pH 7.4) at various microgel concentrations at 37°C. The sample denoted “0 wt % microgel” consists of naltrexone dissolved in PBS in the absence of microgel. The lines are fitted with the aid of Eq. (2).

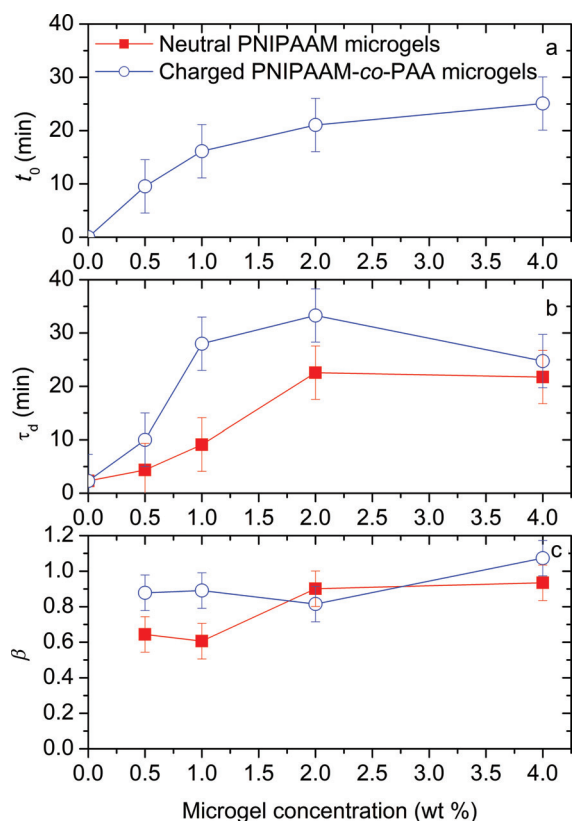


Figure 5. The effect of microgel concentration for the two microgels on (a) the lag time (t_0), (b) the rate constant (τ_d), and (c) the shape parameter (β) for samples containing 0.02 wt % of naltrexone.

Although the neutral microgels did not exhibit any release lag time, a significant lag time was observed for the charged microgels (see Fig. 5a). For most release studies from uncoated micro/nanoparticles, no release lag time is observed, and a burst release effect is much more common.⁴⁸ However, a significant lag time has been detected for some micro- and nanoparticle systems.⁴⁹ In this study, the lag times were attributed to a combination of a very low adsorption of the probe molecules onto the surface of the particles and migration of the probe from the inner regions of the particles to their surface. Accordingly, the lag times observed for the charged microgels indicate that practically all of the naltrexone is encapsulated well within the microgels.

The absence of a lag time for the neutral microgels suggests that the charged groups play a vital role in inducing the lag times. Naltrexone has pK_a values of about 8.1 (proton on N) and 9.5 (phenolic proton) at 37°C,³⁰ and can therefore be expected to be positively charged at the considered pH of 7.4. Considering the relatively high salinity of the buffer, some screening of electrostatic interactions in the system is expected. Unfortunately, the microgel suspensions are too turbid for zeta potential determination at the considered conditions. However, zeta potential measurements of other charged particles at corresponding salinities show that even though the addition of salt reduces the zeta potentials, the particles can still have a significant charge at these conditions.^{50–52} It is therefore reasonable to assume that there are attractive electrostatic interactions between the positively charged drug and the negatively charged PAA groups that are distributed throughout the charged microgels.

Taşdelen et al.¹⁴ found that the addition of charged poly(itaconic acid) groups to PNIPAAm microspheres slowed down the drug release from the particles due to electrostatic interactions with oppositely charged drugs. Although no release lag time was observed for this system, it was found that the charged groups prevented a significant part of the drugs to be released from the particles due to specific binding of the drugs to the dissociated acid groups. In our case, all naltrexone was released from the microgels, but the electrostatic interactions between the drug and the oppositely charged microgels causes a lag time before any drug is released from the particles. The lag time is increasing when the microgel concentration is raised. The reason for this will be discussed below.

Figure 5b shows the rate constant (τ_d), corresponding to the time (after the initial lag time) it takes to release 63.2% of the encapsulated naltrexone (see Eq. 2). It can be observed from Figure 5b that an increase in the microgel concentration results in a slower release of the drug. As increasing the microgel concentration can cause enhanced aggregation of the microgels into larger particles,²⁹ the slower release rates might be due to the larger sizes (see *Evaluation of the Release Rates* above). At high microgel concentrations, t_d is no longer dependent on the amount of microgel present. This could be because the specific leak surface becomes less size dependent for large particles.

It is obvious from Figure 5b that the release of naltrexone is slower for the negatively charged microgels than for the neutral ones. Slower release rates might be caused by larger or more compact particles, or by stronger interactions between the drug and the microgel. However, at very low polymer concentrations, the addition of this small amount of charges to the PNIPAAm microgels had little effect on contraction and aggregation of the microgel suspensions.²⁹ Even if the charges should influence the aggregation and contraction behavior of the microgels at the higher microgel concentrations and higher ionic strength utilized in this study, the charges are expected to reduce the tendency to form large aggregates, and in addition the repulsive electrostatic forces should also make the microgels less compact. It is therefore unlikely that the slower release rates observed for the charged microgels are due to larger particles or a more tight structure within the charged microgels. Accordingly, it is reasonable to assume that the slower release rates are caused by the attractive electrostatic interactions between the drug and the oppositely charged microgels.

The shape parameter, β , is plotted as a function of microgel concentration in Figure 5c. β values in the range 0.39–0.69 indicate diffusion in a fractal or disordered substrate different from a percolation cluster, values in the range 0.69–0.75 suggest diffusion in normal Euclidian space, whereas values between 0.75 and 1 are indicative for diffusion in normal Euclidian space with a contribution of another release mechanism.³⁶ The β values for the charged microgels are all in the range between 0.75 and 1. This seems reasonable as the charges will prevent the microgels from contracting too much; thereby leaving a network that is open enough for the naltrexone to diffuse through as a normal Euclidian space. This β range also indicates that there is an additional release mechanism, which is in agreement with the electrostatic interactions between the drug and the oppositely charged microgels.

At low microgel concentrations, the neutral microgel has β values around 0.69. Accordingly, the release mechanism is either due to diffusion in a fractal or disordered substrate or

in normal Euclidian space. At high microgel concentrations, the β values of the neutral microgels are in the region 0.75–1. This suggests diffusion in normal Euclidian space with a contribution of another release mechanism, which is probably caused by hydrophobic interactions between naltrexone and the microgels.

For both microgel systems, the β values seem to be somewhat higher when the microgel concentration is raised. Increasing the microgel concentration can induce aggregation and thereby the formation of larger particles.²⁹ However, increased particle size is expected to decrease the values of β .⁴² β is also dependent on the intermolecular distances within the matrix, and will increase when the intermolecular distances become larger,⁴² that is, when the matrix structure is more open. This might explain why the charged microgels, which are expected to be more swollen due to electrostatic repulsions, have higher β values than the neutral microgels at low microgel concentrations. It is possible that an increased microgel concentration causes enhanced associations between the polymer chains inside the particles. These associations may lead to the formation of a more inhomogeneous intraparticle network with a more open structure,^{53,54} which might cause the observed increase in β . Another possible explanation for the increased β values at high microgel concentrations is that the release profile is influenced by enhanced hydrophobic interactions between naltrexone and the microgels.

Because the naltrexone seemed to be completely encapsulated within the microgels when a naltrexone concentration of 0.02 wt % was used, it was also interesting to examine whether the microgels were capable of encapsulating even higher amounts of the drug. We therefore carried out some experiments where the microgel concentration was kept constant at 1.0 wt %, and the naltrexone concentration was increased from 0.02 to 0.08 wt %. Figure 6 shows the drug release profiles for the two microgel suspensions at three different naltrexone concentrations. As for the previous results, the curves were fitted to the Weibull equation (Eq. 2), and the fitted curves give a good representation of the measured data. There is a clear change in the release profile when the naltrexone concentration is raised above 0.02 wt % for the charged microgel (Fig. 6b), whereas the effect of naltrexone concentration is less obvious for the neutral microgels (Fig. 6a).

Interestingly, the release lag time of the charged microgels is independent on the naltrexone concentration (see Fig. 7a). We have already established that the lag time is due to electrostatic interactions, and that increasing the microgel concentration causes longer lag times. As the naltrexone concentration does not affect the lag time, the ratio between naltrexone and the charged groups on the microgels is not the determining factor. The microgel concentration dependency (Fig. 5a) therefore indicates that the lag time is dependent on the characteristics of the microgel aggregates. This seems reasonable as the microgels are expected to form larger aggregates when the microgel concentration is increased. The lag time should be dependent on the time it takes the drug to diffuse from its location inside the particles and out to the surface. Therefore, the longer lag times at higher microgel concentrations (Fig. 5a) are probably due to an increased size of the microgel aggregates. At a constant microgel concentration, the sizes of the aggregates do not change much when different amounts of naltrexone are added to the samples, and the lag time is therefore independent of the naltrexone concentration (Fig. 7a).

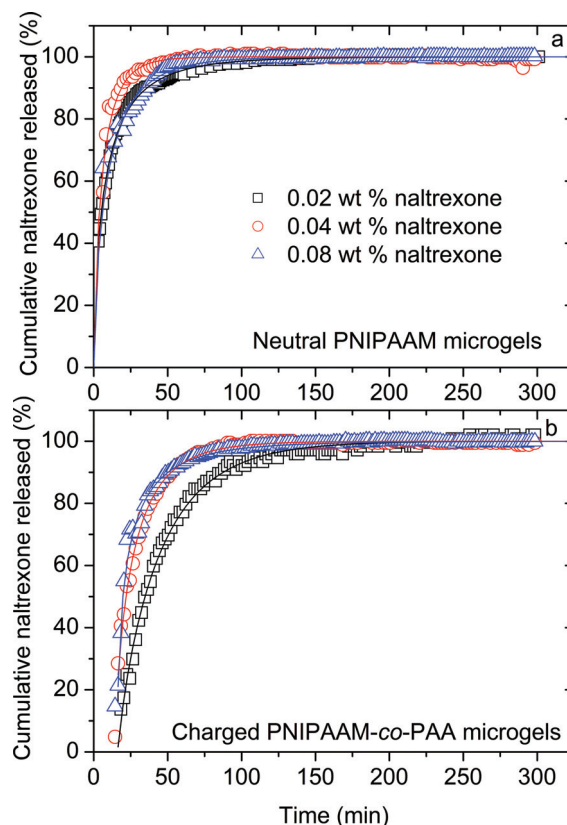


Figure 6. The effect of naltrexone concentration on the *in vitro* release profiles of naltrexone from 1.0 wt % of the two microgels in buffered suspensions (pH 7.4) at 37°C. The lines are fitted with the aid of Eq. (2).

As can be seen from Figure 7b, for the charged microgels the release of the drug is faster when the drug loading is raised. However, for the neutral microgels there is no significant effect of naltrexone concentration on the release rates. At the highest naltrexone concentration, the release times for the two microgels are quite close to each other. This suggests that the faster release rates for the charged microgels at higher naltrexone concentrations might be due to a reduction of available charged groups on the microgels per naltrexone unit.

For the neutral microgels, the shape parameter, β , is not significantly affected by the naltrexone concentration. However, for the charged microgels, the value of β decreases as the naltrexone concentration is raised (see Fig. 7c). It is possible that more of the charged PAA groups are associated with the oppositely charged drug at high naltrexone concentrations. Accordingly, the effective amount of repulsive forces within the microgels decreases. Without the repulsive forces, a more compact network is formed within the microgels (reminiscent of the neutral microgels). This is in agreement with Monte Carlo simulations, which predict a decrease of β when the intermolecular distances within the matrix decreases.⁴²

CONCLUSIONS

Poly(*N*-isopropylacrylamide) microgels without attached PAA chains and negatively charged PNIPAAm-co-PAA microgels were found to show sustained release of naltrexone.

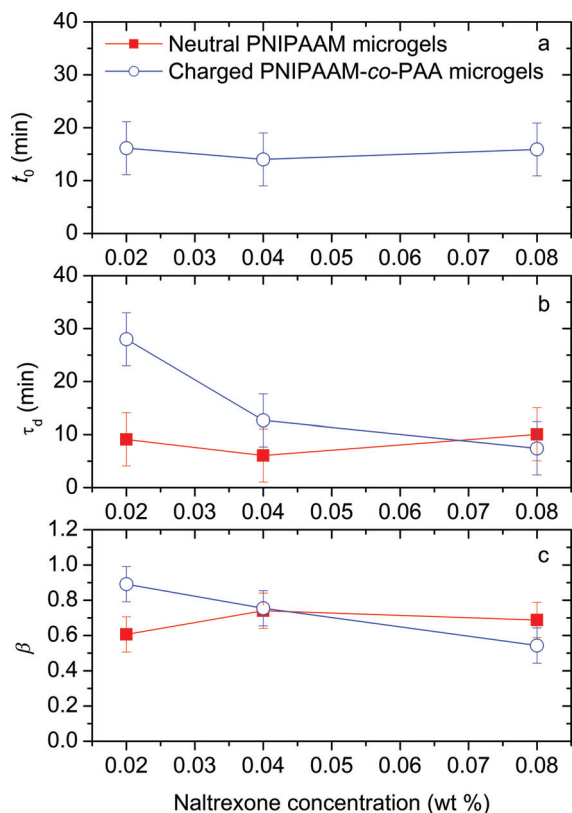


Figure 7. The effect of naltrexone concentration for the two microgels on (a) the lag time (t_0), (b) the rate constant (τ_d), and (c) the shape parameter (β) for samples containing 1.0 wt % of the two microgels.

The release curves could successfully be represented by the Weibull equation.

Attractive associative interactions between naltrexone and oppositely charged microgels cause slower release times compared with the neutral microgels. The charges also cause the drug to diffuse further into the microgel aggregates, resulting in a release lag time.

Increasing the polymer concentration is expected to cause enhanced aggregation of the microgel suspensions. The resulting aggregates have a smaller specific leak surface, and accordingly the drug release times become longer. The larger aggregates also increase the release lag time of the charged microgels.

For the neutral microgels, the drug release is not significantly influenced by the naltrexone concentration. However, for the charged microgels both the drug release rates and the β values are declining when the naltrexone content becomes higher. This might be caused by the reduced ratio between the charged groups on the microgels and the oppositely charged naltrexone. However, the amount of charged groups on the microgels is still high enough to prevent the release lag time from being affected by the naltrexone concentration.

The tested concentrations of the microgels did not exhibit any significant cytotoxic effect on a HeLa cells, making them promising candidates for drug delivery purposes. They also have the additional advantage of providing a system where the drug is easily loaded into the microgels, without any time-consuming preparation procedures or the use of harsh solvent conditions or severe mechanical stress.

ACKNOWLEDGMENT

Authors gratefully acknowledge financial support from the Norwegian Research Council for the project 190403.

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